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Patent and Trade Mark Attorneys

17 September 2002

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**SPECIAL DELIVERY
RETURN RECEIPT REQUESTED**

Dear Ian

**US Patent Application No 10/081,735
Fuel Composition**

I refer to Alan Rae's letter to you of 21 February 2000.

The abovementioned US Patent application, serial No. 10 081,735, is a continuation of US Patent Application Serial No. 09/868,420, which is itself based on International Patent application No. PCT/GB99/04155 (publication No. WO 00/36055). As you are aware you are named as a co-inventor on this Patent application with Ian Williamson.

Since the necessary documentation in support of the abovementioned application is still outstanding, I am enclosing herewith (ribboned to this letter) the following documentation and would request that you execute the appropriate documents personally where indicated and return all of the documents to me in the stamped addressed envelope provided, by no later than **1 October 2002**.

The documents enclosed are:

1. Copy of published PCT Patent application, including bibliographic details, specification, claims and drawings.
2. Combined Declaration and Power of Attorney re 10 081,735.

A corresponding set of documents has been sent to Ian Williamson as co-inventor.

We warrant the name and professional qualifications of the person
employed to execute this letter as above.

Yours faithfully,
Harrison Goddard Foote

By: [Signature]

Enclosed for Ian Williamson: 1 copy of the above documents

harrison goddard foote

[illegible]

Charles Hewetson - Richards Butler, Solicitors (w:en:es)

COMBINED DECLARATION FOR PATENT APPLICATION AND POWER OF ATTORNEY

(Includes Reference to PCT International Applications)

ATTORNEY DOCKET NUMBER
09520J489US0

As a below named inventor, I hereby declare that:

My residence, post office address and citizenship are as stated below next to my name.

I believe I am the original, first and sole inventor (if only one name is listed below) or an original, first and joint inventor (if plural names are listed below) of the subject matter which is claimed for and which a patent is sought on the invention entitled:

FUEL COMPOSITION

the specification of which (check only one item below):

☐ is attached hereto.☐ was filed as United States application

Serial No. _____

on _____

and was amended

on _____ (if applicable).

☒ was filed as PCT international applicationNumber PCT/GB99/04155on 15 December 1999

and was amended under PCT Article 19

on _____ (if applicable).

I hereby state that I have reviewed and understand the contents of the above-identified specification, including the claims as amended by any amendment referred to above.

I acknowledge the duty to disclose information which is material to the examination of this application in accordance with Title 37, Code of Federal Regulations, §1.56(a).

I hereby claim foreign priority benefits under Title 35, United States Code, §119 of any foreign application(s) for patent or inventor's certificate or of any PCT international application(s) designating at least one country other than the United States of America listed below and have also identified below any foreign application(s) for patent or inventor's certificate or any PCT international application(s) designating at least one country other than the United States of America filed by me on the same subject matter having a filing date before that of the application(s) of which priority is claimed:

PRIOR FOREIGN/PCT APPLICATION(S) AND ANY PRIORITY CLAIMS UNDER 35 U.S.C. 119:

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Great Britain	9827592.8	15 December 1998	<input checked="" type="checkbox"/> YES	<input type="checkbox"/> NO
			<input type="checkbox"/> YES	<input type="checkbox"/> NO
			<input type="checkbox"/> YES	<input type="checkbox"/> NO
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US

Combined Declaration for Patent Application and Power of Attorney (Continued) (Includes Reference to PCT International Applications)				ATTY'S DOCKET NUMBER 0952/0J499US0	
<p>I hereby claim the benefit under Title 35, United States Code, §120 of any United States application(s) or PCT international application(s) designating the United States of America that is/are listed below and, insofar as the subject matter of each of the claims of this application is not disclosed in that those prior application(s), in the manner provided by the first paragraph of Title 35, United States Code §112, I acknowledge the duty to disclose material information as defined in Title 37, Code of Federal Regulations, §1.56(a) which occurred between the filing date of the prior application(s) and the national or PCT international filing date of this application:</p>					
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PCT APPLICATION NO.	PCT FILING DATE	U.S. SERIAL NUMBER ASSIGNED (if any)			
PCT/GB99/04185	15 December 1999				
POWER OF ATTORNEY: As a named inventor, I hereby appoint the following attorney(s) and/or agents to prosecute this application and transact all business in the Patent and Trademark Office connected therewith. Morris Relson #15,108, Gordon D. Coplein #19,165, William F. Dudine, Jr. #20,569, Michael J. Sweedler #19,937, S. Peter Ludwig #25,351, Paul Fields #20,298, Joseph B. Lerch #26,936, Melvin C. Garner #26,272, Ethan Horwitz #27,646, Beverly B. Goodwin #28,417, Adda C. Gogoris #29,714, Martin E. Goldstein #20,869, Bert J. Lewen #19,407, Henry Sternberg #22,408, Peter C. Schechter #31,662, Robert Schaffer #31,194, Robert C. Sullivan, Jr. #30,499, and Joseph R. Robinson #33,448					
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2 0 3	FULL NAME OF INVENTOR	FAMILY NAME	FIRST GIVEN NAME	SECOND GIVEN NAME	
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<p>I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true, and further that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under section 1001 of Title 18 of the United States Code, and that such willful false statements may jeopardize the validity of the application or any patents issuing thereon.</p>					
SIGNATURE OF INVENTOR 201		SIGNATURE OF INVENTOR 202		SIGNATURE OF INVENTOR 203	
DATE		DATE		DATE	

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INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

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(21) International Application Number: PCT/GB99/04155 (22) International Filing Date: 15 December 1999 (15.12.99) (30) Priority Data: 9827592.8 15 December 1998 (15.12.98) GB (71) Applicant (for all designated States except US): AAE HOLDINGS PLC [GB/GB]; Unit 11, Bridge Road, Business Park, Haywards Heath, West Sussex RH16 1TX (GB). (72) Inventors; and (75) Inventors/Applicants (for US only): HAZEL, Clifford, James [GB/GB]; Keveral Mill, Hessenford, Cornwall PL11 3HW (GB). WILLIAMSON, Ian, Vernon [GB/GB]; 27 Pownall Avenue, Bramhall, Stockport, Cheshire SK7 2HE (GB). (74) Agent: GILHOLM, Steve; Harrison Goddard Foote, Belmont House, 20 Wood Lane, Leeds LS6 2AE (GB).	(81) Designated States: AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZA, ZW, ARIPO patent (GH, GM, KE, LS, MW, SD, SL, SZ, TZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG). Published <i>With international search report.</i> <i>Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</i>	
(54) Title: FUEL COMPOSITION		
(57) Abstract <p>There is described a method of reducing the vapour pressure (RVP) of a gasoline/alcohol mixture which comprises adding a surfactant composition comprising an alkanolamide, an alkoxylated alcohol and an alkoxylated fatty acid to a gasoline/alcohol mixture.</p>		

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FUEL COMPOSITION

The invention relates to a novel method of reducing the vapour pressure of fuels and to the new use of compositions.

5

It is well known that liquid fuels when burned in an internal combustion engine can give rise to pollution and other undesired side effects. Furthermore, the relatively high vapour pressure of most liquid hydrocarbon fuels is known to give rise to environmental and safety problems upon storage and transportation. Numerous proposals have been advanced to reduce these side effects and enhance efficiency, eg 10 miles per gallon. It has been realised that surfactants can play a useful role in this context but so far as we are aware none has satisfied the modern commercial criteria. It is one object of this invention to meet the need.

15 According to the US Environmental Protection Agency (EPA), volatile organic compound (VOC) emissions from gasoline are, as ozone precursors, a major contributor to the nation's serious ground level ozone problem, which harms human health and the public welfare. In addition to automotive emissions, gasoline evaporation occurs during refining, during commercial handling, transportation, and 20 storage, and during refuelling. Because gasoline evaporation is a significant environmental problem, work was initiated at the National Alternative Fuels Laboratory (NAFL) of the University of North Dakota Energy & Environmental Research Center (EERC) to investigate evaporation rates of various gasolines and gasoline-oxygenate blends and the composition of their evaporative emissions. One 25 aspect of the work, which was funded by the US Department of Agriculture (USDA), was to compare evaporation rates and compositions of E10 evaporative emissions with those of nonethanol-blended base fuels.

Widespread commercial use of E10 fuels was originally promoted as a means to 30 reduce gasoline consumption during the oil embargo of 1973. Current interest in ethanol as a transportation fuel involves its use as an oxygenate to reduce carbon

monoxide (CO) combustion emissions. Also, its octane-boosting quality makes it useful as a replacement for octane-supplying aromatics that are slated for removal from reformulated gasolines to be sold in ozone non-attainment areas. Despite its desirable qualities as a fuel and the fact that it is renewable, ethanol has serious
5 competition. Ethanol supporters are locked in a long-term debate with supporters of the other major commercially available gasoline oxygenate, methyl *tertiary*-butyl ether (MTBE), over which oxygenate is best for the environment.

A major point of contention in the debate is the approximate one-pound per square
10 inch (psi) increase in RVP that occurs upon addition of ethanol to base fuel at a concentration of 10 vol%. However, while RVP is a useful indicator of a fuel's volatility, it is less useful as an indicator of how atmospheric interaction of vapour from the fuel (generated via evaporation or displacement during tank filling) may affect the environment in terms of ground level ozone production. Because gasolines
15 are complex mixtures of over 200 components, the overall ozone-producing potential of a particular gasoline's vapour will depend not only on how much vapour is emitted, but also on the ozone-producing potential and concentration of each vapour component.

20 Surfactants have long had utility as additives which can affect the performance of fuels such as gasoline and diesel. British Patent No 2217229 describes a solubilising compound as a fuel additive. In particular there is described a composition comprising:

- 48 parts by volume of an alcohol ethoxylate;
- 25 3-8 parts by volume of lauric diethanolamide;
- 3-8 parts by volume of oleic diethanolamide; and
- 1.5-4 parts by volume of a polyglycol ether such as ethoxylated oleic acid.

Such compositions are useful as fuel additives and enable the solubilisation of water
30 in fuels thus reducing its potential corrosive effect. However, the compositions are disadvantageous in that, *inter alia*, they require a high additive to fuel ratio.

Furthermore, they do not address the problems of emissions of gases such as CO, CO₂ and NOX, nor do they address the problems of evaporative loss due to the high vapour pressure of the fuel composition.

- 5 International Patent Application No WO 98/17745 describes an alternative formulation which comprises.

25% w/w of a diethanolamide,

50% w/w of an ethoxylated alcohol, and

- 10 25% w/w of a fourteen carbon chain fatty acid with seven ethoxylate groups.

WO '745 especially describes fuel compositions comprising, *inter alia*, an additive made up of a fatty acid diethanolamide, an alcohol ethoxylate and an ethoxylate of a fatty acid, the degree of ethoxylation being selected so that a long term stable fuel
15 composition is formed and, in particular, wherein, by carefully selecting the degree of ethoxylation, a balanced blend can be produced, such that the molecular weight ratios of each of the three components are substantially equivalent, for example, the molecular weight ratios are 1:1:1.

- 20 Whilst such additives provide significant reductions in emissions and are useable at low concentrations, there is still a need for a fuel composition which is capable of reducing emissions whilst maintaining performance levels.

It is known that to meet some legislation requirements fuels must have a low Reid
25 Vapour Pressure value (RVP). e.g. under California State law, the RVP must be below 7psi.

Some modern day gasoline contains a small percentage of a condensate from natural gas pipelines. Gasoline is made up from a range of hydrocarbons, which is
30 determined by the refining and blending processes. They may also contain minor amounts of comparatively volatile contaminants picked up during transportation, for

example, through pipe lines which have been previously used for transportation of natural gases. Pentanes are a range of C5 hydrocarbons that are typically left in the crude oil cut after the gas condensates are removed. Natural gas (methane) is extracted first from the oil field, ethane is increasingly also extracted at the well for
5 petrochemical production of ethylene, while LPG, a mixture of C3 and C4 hydrocarbons, is liquefied under moderate pressure for use as a gas fuel. Pentane is usually the first cut off the Crude Distillation Unit at the refinery and the disposal of pentane creates a problem for the petroleum industry. Pentane has little value as a petrochemical feedstock, and blending it into the gasoline pool has the undesirable
10 effect of raising the RVP. The aforementioned contaminants, as well as low boiling fractions arising from the refining and blending processes, all have the undesirable effect of increasing the vapour pressure of the gasoline. This undesirable effect will be further exacerbated if the fuel is blended with certain volatile oxygenating agents such as ethanol. Thus the condensate is a light fraction which is primarily pentanes
15 although other hydrocarbons may be present. The condensate adversely affects the RVP, especially when the gasoline contains an alcohol such as ethanol.

It is known that when gasoline and an alcohol are mixed, the resultant mixture has an increased vapour pressure. We have now surprisingly found that by adding certain
20 surfactant compositions to a gasoline and alcohol mixture a reduction in vapour pressure (RVP) can be achieved.

Thus, according to the invention we provide a method of reducing the vapour pressure of a gasoline/alcohol mixture which comprises adding a surfactant
25 composition comprising an alkanolamide, an alkoxyated alcohol and an alkoxyated fatty acid to a gasoline/alcohol mixture.

By the term vapour pressure we mean the Reid Vapour Pressure (RVP). Thus the method of the invention preferably reduces the RVP of the gasoline mixture to less
30 than 8psi, preferably less than 7psi and especially between 6 and 7psi.

For the purposes of this invention the RVP is measured using the test method of ASTM D5191. Thus the vapour pressure is given as psi at 37.8°C.

5 It is a novel aspect of the invention to use the aforementioned surfactant composition in the reduction of RVP.

According to a further aspect of the invention we provide the use of an additive composition as hereinbefore described in the manufacture of a fuel composition having an RVP of less than 8psi.

10

The alkanolamide is preferably an ethanolamide and more preferably a diethanolamide. Especially preferred are the diethanolamides and particularly the super diethanolamides. By the term super diethanolamide we mean a diethanolamide in which the nitrogen is substituted by an acyl substituent e.g. acyl C₅ to C₂₀,
15 preferably C₈ to C₁₆, more preferably C₁₀ to C₁₅. The most preferred diethanolamide has a C₁₂ acyl substituent i.e. lauryl diethanolamide. Where the amides are derived from natural products this moiety will have an even number of carbon atoms, e.g. 12 for the lauryl derivative. Note, the alkyl part of this group is the R group which will be an odd number.

20

There are three commercial routes to alkanolamides;

Acid + alkanolamine = alkanolamide + water

Plant or animal oil (triglyceride) + alkanolamine = alkanolamide + glycerol

25

Methyl ester + alkanolamine = alkanolamide + methanol

These are listed in order of increasing product quality. The route via the acid often uses an excess of alkanolamine to produce a product higher in amide than is obtainable from the acid if a stoichiometric ratio is used; these products are
30 sometimes referred to as super amides.

The alkoxyated alcohol is preferably an ethoxylated alcohol. It is essential that the ethoxylated alcohol is an oil soluble alcohol. Therefore, alkanols are preferred and these may be primary, secondary or tertiary alkanols and especially primary alkanols. As the oil solubility of the alcohol may vary with the carbon chain length of the ethoxylated alkanol, the alkanol is preferably a C₅ to C₂₂ alkanol, more preferably C₅ to C₁₅ alkanol. The ethoxylated alcohol may comprise a mixture of alkanols. However, it is preferred that in such mixtures one alkanol will predominate. Thus, the most preferred alkanol is predominantly a C₉ to C₁₁ alkanol. In addition the degree of ethoxylation of the alcohol may be varied and the oil solubility will, generally, decrease with the increase in the degree of ethoxylation. It is preferred that the ethoxylate to alcohol ratio is greater than 2. More preferably, the ethoxylate to alcohol ratio is from between 1 and 10, preferably between 1 and 5, more preferably between 1 and 3 and especially between 2 and 3. A commercially available ethoxylated alcohol is especially preferred in which the ethoxylate to alcohol ratio is 2.75. Such an alcohol ethoxylate is available as NEODOL 91/2.5.

The fatty acid ethoxylate may comprise any conventionally known fatty acid ethoxylate. Thus the fatty acid ethoxylate may be derived from a fatty acid having from 8 to 20 carbon atoms, preferably from 10 to 18 and most preferably 14 carbon atoms (myristic acid).

The degree of ethoxylation is chosen to optimise performance in the blend with the other two selected surfactants and may be from 1 to 20, but more preferably from 5 to 12. A suitable product within this range would be, for example that derived from the addition of 7 molecules of ethylene oxide to 1 mole of myristic acid.

The values will depend on the length of the hydrophilic chain, typically an ethoxylate chain. The length of the chain will increase the extent of solubilisation because of a greater ability to solubilise.

As with the compositions described in WO98 17745, a blend of surfactants is preferred, preferably by selecting one appropriate to the fuel, say 10 to 18 for hydrocarbon fuel, most preferably 13. In the case of an alcohol the HLB value of the surfactant is between 3 and 7, most preferably about 4. But the addition of
5 surfactants normally create ratios of 1:1 or high volume emulsions or 5:1 ratios when the solubalisation is required at 1:100.

The invention has the ability to unify the HLB requirements of any liquid fuel which in turn allows for one dose to be used in any fuel from C5 carbon chains up. The
10 benefit being the amount of treatment directly related to the co-solvency ability.

Preferably the ethoxylate of the fatty acid makes up about 25% by volume of the additive and further preferably the alcohol ethoxylate comprises 50% by volume of the additive.
15

The surfactant additive may be added to a hydrocarbon fuel, eg diesel, petrol or alcohol, such as ethanol which may or may not be contaminated with water.

The fuel composition comprising a fuel and a fuel additive wherein the additive to
20 fuel ratio is 0.5 – 1:1200.

The alcohol to hydrocarbon fuel ratio is from 1:8 to 1:9. Water may be present, e.g. as condensate, or added separately, in water: alcohol: fuel ratio of 0:1:8 to 0.5:1:9 and preferably from 0.25:1:8 to 0.25:1:9.
25

The additive may be present in an additive to fuel/alcohol ratio of about 0.5 – 1:1000. Preferably the additive to fuel ratio is about 1:1000, most preferably about 1:1200. In this context the fuel is the hydrocarbon fuel and alcohol mixture.

Example 1

Different blends shown in Table 1 were made and subjected to gas chromatography. The ratio of hydrocarbon: ethanol was measured by gas chromatography. The results were used to indicate the extent to which the content of the ethanol could be incorporated without exceeding an increase in the temperature at which the volatile ingredients evaporated. The results showed the amount of ethanol that could be incorporated into the hydrocarbon without increasing the temperature at which the volatile components evaporated.

The additive comprised 60% by weight ethanolethoxylate, 20% by weight diethanolamide and 20% by weight of lauric diethanolamide. Our investigation suggests that these interact to form a blend in which the ethanolamide is a backbone.

Example 2

Different blends were made up and the RVP determined. The following results were obtained:

Table II

Sample	Gasoline	Ethanol	Additive as above	Water	RVP (1)	RVP (2)	RVP (3)
1	90	10	-	-	7.48	7.7	7.58
2	87.5	10	2.5	-	7.18	7.41	7.29
3	85.0	10	2.5	2.5	6.87	7.11	6.98
4	85.0	10	2.5	2.5	6.96	7.19	7.06

The RVP value was measured according to ASTM D5191 and is the mean value of results calculated according to CARB (1), EPA (2) and ASTM (3) methods. In Samples 2, 3 and 4 the gasoline included a proportion of a condensate which was mainly pentanes. These results demonstrate a tendency of the additive to reduce the

RVP value, and that by selecting appropriate proportions a fuel composition can be made to meet the requirements of local legislation.

Table 1

Sample	Hydrocarbon (a)	Ethanol (b)	Pentane	Additive	Ratio (a):(b)
1	90	10	0	0	15:1
2	89	10	0	1	16.7:1
3	82	13.5	4.5	0	17.75:1
4	77.5	13.5	4.5	4.5	11.69:1

5

Example 3

Various tests were conducted according to ASTM D-5191. The results are given in
10 Tables III to IX.

Example 4

15 Various tests were conducted according to the 1975 US Federal City Gasoline Test.
The results are given in Tables X to XIII:

Table X: Tests on CARB Gasoline

Table XI: Tests on EPA

20 Table XII: Tests on Gasohol

25

CLAIMS

1. A method of reducing the vapour pressure of a gasoline/alcohol mixture which comprises adding surfactant composition comprising an alkanolamide, an
5 alkoxyated alcohol and an alkoxyated fatty acid to a gasoline/alcohol mixture.
2. A method according to Claim 1 characterised in that the RVP is less than 7psi.
- 10 3. A method according to Claim 2 characterised in that the RVP is less than 7psi.
4. A surfactant composition according to Claim 3 characterised in that the RVP is between 6 and 7psi.
- 15 5. A method according to Claim 1 characterised in that the alkanolamide is a diethanolamide.
6. A method according to Claim 5 characterised in that the nitrogen in the
20 diethanolamide is substituted by an alkyl C₅ to C₂₀ substituent.
7. A surfactant composition according to Claim 6 characterised in that the diethanolamide is a lauryl diethanolamide.
- 25 8. A method according to Claim 1 characterised in that the alkoxyated alcohol is an ethoxyated alcohol.
9. A method according to Claim 8 characterised in that the ethoxyated alcohol is a C₅ to C₁₇ alkanol.

30

10. A method according to Claim 8 characterised in that the ethoxylated alcohol comprises a mixture of alkanols in which one alkanol predominates.
11. A method according to Claim 8 characterised in that the predominate alkanol
5 is a C₆ to C₁₁ alkanol.
12. A method according to Claim 8 characterised in that the ethoxylate to alcohol ratio is from between 1 and 5.
- 10 13. A method according to Claim 8 characterised in that the ethoxylated alcohol is NEODOL 91/2.5.
14. A method according to Claim 1 characterised in that the fatty acid group is a C₈ to C₂₀ fatty acid.
15
15. A method according to Claim 14 characterised in that the fatty acid group is a C₁₄ fatty acid (myristic acid).
16. A method according to Claim 1 characterised in that the ester moiety of the
20 fatty acid ester is an alkyl ester.
17. A method according to Claim 16 characterised in that the alkyl group is a C₁ to C₁₀ alkyl.
- 25 18. A method according to Claim 1 characterised in that the composition comprises 25% v/v of the fatty acid ester.
19. A method according to Claim 1 characterised in that the composition comprises 50% v/v of the alcohol ethoxylate.
30

20. A method according to claim 1 characterised in that the surfactant additive to fuel/alcohol ratio is from 0.5:1200 to 1:1000.

21. The use of a surfactant composition comprising an alkanolamide, an
5 alkoxyated alcohol and an alkoxyated fatty acid ester in the manufacture of a fuel composition having an RVP of less than 8psi.

22. A method of manufacturing a mixture comprising gasoline, alcohol and a
surfactant composition, said surfactant composition comprising an alkanolamide, an
10 ethoxyated alcohol and an alkoxyated fatty acid characterised in that the method comprises blending the alcohol and surfactant followed by blending with gasoline.

23. A method substantially as described with reference to the accompanying
examples.
15

20

25

1/10

TABLE III

LABORATORY TESTS RESULTS				
05/28/99				
JOB NUMBER: 990554		ATTN: Attn		
CLIENT I.D.....	C - BASE CARBOB (GASOLINE)	LABORATORY I.D.....	990554-0034	
DATE SAMPLED.....	05/26/99	DATE RECEIVED.....	05/26/99	
TIME SAMPLED.....	12:02	TIME RECEIVED.....	12:02	
WORK DESCRIPTION.....	C	REMARKS.....		

TEST DESCRIPTION	FINAL RESULT	LIMITS* DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/26/99
RVP CARB CALCULATION	5.82*	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	6.08	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	5.94	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	6.73	0.10	PSI	ASTM D-5191	

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TABLE IV

LABORATORY TESTS RESULTS					
05/28/99		ATTN: Attn			
JOB NUMBER: 990554					
CLIENT I.D.....	CARBOB 90/10	BASE + 10% ETOH			
DATE SAMPLED.....	05/20/99				
TIME SAMPLED.....	16:01				
WORK DESCRIPTION....	CARBOB 90/10				
LABORATORY I.D.....	990554-0039				
DATE RECEIVED.....	05/20/99				
TIME RECEIVED.....	16:01				
REMARKS.....	1-500ml amber btl				
TEST DESCRIPTION	FINAL RESULT	LIMITS/DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	7.16*	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	7.41	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	7.29	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	8.12	0.10	PSI	ASTM D-5191	

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TABLE V

LABORATORY TESTS RESULTS					
05/28/99		ATTN: Attn			
JOB NUMBER: 990554					
CLIENT I.D.....	CARBOB 90/10 LN	BASE + 10% ETOH	LABORATORY I.D.....	990554-0040	
DATE SAMPLED.....	05/20/99	+ 1% 25% NEODOL	DATE RECEIVED.....	05/20/99	
TIME SAMPLED.....	16:01	37.5% OPE	TIME RECEIVED.....	16:01	
WORK DESCRIPTION.....	CARBOB 90/10 LN	37.5% ATLAS	REMARKS.....	1-500ml amber btl	
TEST DESCRIPTION	FINAL RESULT	LIMITS*/DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	6.79	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	7.04	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	6.91	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	7.73	0.10	PSI	ASTM D-5191	

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TABLE VI

LABORATORY TESTS RESULTS					
05/28/99		ATTN: Attn			
JOB NUMBER: 990554					
CLIENT I.D.....	CARBOB 90/10 REVERSE	BASE + 10% ETOH	LABORATORY I.D.....	990554-0042	
DATE SAMPLED.....	05/20/99	+ 1% AAE + 75% NEODOL	DATE RECEIVED.....	05/20/99	
TIME SAMPLED.....	16:01	+ 12.5% ODE	TIME RECEIVED.....	16:01	
WORK DESCRIPTION....	CARBOB 90/10 REVERSE	+ 12.5% ATLAS	REMARKS.....	1-500ml amber btl	
TEST DESCRIPTION	FINAL RESULT	LIMITS*/DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	6.56	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	6.80	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	6.67	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	7.48	0.10	PSI	ASTM D-5191	

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TABLE VII

LABORATORY TESTS RESULTS				
JOB NUMBER: 990554		05/28/99		
CLIENT I.D.: CARBOB 90/10 (RECHECK)		ATTN: Attn		
DATE SAMPLED: 05/27/99		LABORATORY I.D.: 990554-0043		
TIME SAMPLED: 16:01		DATE RECEIVED: 05/27/99		
WORK DESCRIPTION: CARBOB 90/10 (RECHECK)		TIME RECEIVED: 16:01		
		REMARKS: 1-500ml amber btl		
TEST DESCRIPTION	FINAL RESULT	LIMITS* DILUTION	UNITS OF MEASURE	TEST METHOD
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191
RVP CARB CALCULATION	7.22	0.10	PSI	ASTM D-5191 (CARB)
RVP EPA CALCULATION	7.47	0.10	PSI	ASTM D-5191 (EPA)
RVP ASTM CALCULATION	7.34	0.10	PSI	ASTM D-5191 (ASTM)
PRESSURE TOTAL/GRABNER INSTRUMENT	8.18	0.10	PSI	ASTM D-5191

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TABLE VIII

LABORATORY TESTS RESULTS				
05/28/99				
JOB NUMBER: 990554		ATTN: Attn		
CLIENT I.D.	CARBOB 90/10 LN (RECHECK)	LABORATORY I.D.	990554-0044	
DATE SAMPLED	05/27/99	DATE RECEIVED	05/27/99	
TIME SAMPLED	16:01	TIME RECEIVED	16:01	
WORK DESCRIPTION	CARBOB 90/10 LN (RECHECK)	REMARKS	1-500ml anber btl	

TEST DESCRIPTION	FINAL RESULT	LIMITS* DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	6.76	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	7.01	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	6.88	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	7.70	0.10	PSI	ASTM D-5191	

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TABLE IX

LABORATORY TESTS RESULTS					
JOB NUMBER: 990554		05/28/99			
CLIENT I.D.: CARBOB 90/10 REVERSE (RECHECK)		ATTN: Attn			
DATE SAMPLED.....: 05/27/99		LABORATORY I.D.....: 990554-0046			
TIME SAMPLED.....: 16:01		DATE RECEIVED.....: 05/27/99			
WORK DESCRIPTION....: CARBOB 90/10 REVERSE (RECHECK)		TIME RECEIVED.....: 16:01			
		REMARKS.....: 1-500ml amber btl			
TEST DESCRIPTION	FINAL RESULT	LIMITS* DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	6.56	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	6.80	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	6.67	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	7.48	0.10	PSI	ASTM D-5191	

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TABLE X

Fuel Summary CARB Gasoline vs CARB Gasoline + AAE10 @ 1000ppm

Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	Nox g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5736	1993 Ford Explorer	CARB Base, Seq#1	0.151	1.697	0.239	524.16	0.027	0.1196	16.183
5736	1996 GMC Jimmy	CARB Base, Seq#1	0.218	2.744	0.288	485.79	0.03	0.1835	17.387
5675	1998 Dodge Ram	CARB Base, Seq#1	0.135	1.427	0.132	582.37	0.021	0.1108	14.586
5682	1991 Geo Prizm	CARB Base, Seq#1	0.348	3.269	0.372	289.68	0.031	0.3129	28.84
5736	1996 Nissan Sentra	CARB Base, Seq#1	0.073	0.772	0.148	289.58	0.008	0.0638	29.323
5734	1994 Nissan Sentra	CARB Base, Seq#1	0.17	1.878	0.368	308.26	0.017	0.1511	27.373
		Average	0.183	1.964	0.258	413.305	0.022	0.157	22.282
Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5691	1993 Ford Explorer	CARB + AAE10, Seq#2	0.18	1.566	0.245	532.46	0.029	0.1479	15.935
5684	1996 GMC Jimmy	CARB + AAE10, Seq#2	0.202	2.524	0.287	493.70	0.028	0.1695	17.125
5686	1998 Dodge Ram	CARB + AAE10, Seq#2	0.194	1.712	0.115	592.59	0.024	0.1662	14.32
5700	1991 Geo Prizm	CARB + AAE10, Seq#2	0.326	3.416	0.41	287.52	0.036	0.2846	29.035
5745	1996 Nissan Sentra	CARB + AAE10, Seq#2	0.074	0.785	0.129	294.62	0.008	0.0659	28.82
5747	1994 Nissan Sentra	CARB + AAE10, Seq#2	0.182	1.869	0.382	308.29	0.018	0.1613	27.369
		Average	0.193	1.982	0.263	418.199	0.024	0.166	22.101
		Deviation +/-	5.66%	0.89%	1.92%	1.18%	7.51%	5.70%	-0.81%

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TABLE XI

Fuel Summary EPA vs EPA + AAE10 @ 1000ppm

Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5706	1993 Ford Explorer	EPA Base, Seq#3	0.223	2.316	0.34	539.82	0.043	0.1743	15.682
5693	1996 GMC Jimmy	EPA Base, Seq#3	0.242	2.882	0.453	499.87	0.042	0.1947	16.893
5692	1998 Dodge Ram	EPA Base, Seq#3	0.249	2.466	0.19	600.43	0.046	0.1965	14.103
5708	1991 Geo Prizm	EPA Base, Seq#3	0.327	3.392	0.446	294.90	0.035	0.2878	28.328
5752	1996 Nissan Sentra	EPA Base, Seq#3	0.165	1.038	0.208	299.19	0.012	0.1514	28.318
5753	1994 Nissan Sentra	EPA Base, Seq#3	0.253	2.952	0.333	315.84	0.023	0.2271	26.562
		Average	0.243	2.508	0.328	425.009	0.034	0.205	21.648
Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5709	1993 Ford Explorer	EPA + AAE10, Seq#4	0.234	2.744	0.351	534.78	0.048	0.1804	15.808
5701	1996 GMC Jimmy	EPA + AAE10, Seq#4	0.324	3.141	0.517	500.54	0.046	0.2711	16.848
5702	1998 Dodge Ram	EPA + AAE10, Seq#4	0.315	3.308	0.227	598.06	0.056	0.2521	14.123
5717	1991 Geo Prizm	EPA + AAE10, Seq#4	0.365	3.952	0.519	294.89	0.034	0.3268	28.235
5758	1996 Nissan Sentra	EPA + AAE10, Seq#4	0.185	1.299	0.187	300.57	0.014	0.1697	28.146
5759	1994 Nissan Sentra	EPA + AAE10, Seq#4	0.30	3.308	0.371	313.13	0.025	0.2712	26.729
		Average	0.287	2.959	0.362	423.663	0.037	0.245	21.648
		Deviation +/-	18.00%	17.99%	10.27%	-0.32%	10.16%	19.45%	0.00%

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TABLE XII

Fuel Summary Gasohol vs Gasohol + AAE00 @ 4600ppm

Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	Nox g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5722	1993 Ford Explorer	Gasohol Base, Seq#5	0.165	1.545	0.251	519.43	0.033	0.1275	16.336
5711	1996 GMC Jimmy	Gasohol Base, Seq#5	0.219	2.359	0.342	495.15	0.037	0.1765	17.083
5710	1998 Dodge Ram	Gasohol Base, Seq#5	0.219	2.102	0.194	589.97	0.047	0.1655	14.367
5727	1991 Geo Prizm	Gasohol Base, Seq#5	0.273	3.804	0.452	283.54	0.037	0.2309	29.549
5764	1996 Nissan Sentra	Gasohol Base, Seq#5	0.118	0.907	0.175	295.48	0.012	0.1046	28.706
5766	1994 Nissan Sentra	Gasohol Base, Seq#5	0.194	2.237	0.388	311.28	0.022	0.1694	27.056
	Average		0.198	1.992	0.30	415.807	0.032	0.162	22.183
Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5726	1993 Ford Explorer	Gasohol + AAE00, Seq#6	0.17	1.48	0.26	515.20	0.035	0.1312	16.472
5770	1996 GMC Jimmy	Gasohol + AAE00, Seq#6	0.231	1.972	0.30	488.00	0.033	0.1938	17.351
5718	1998 Dodge Ram	Gasohol + AAE00, Seq#6	0.223	2.04	0.156	590.74	0.044	0.1733	14.351
5735	1991 Geo Prizm	Gasohol + AAE00, Seq#6	0.283	2.478	0.443	280.95	0.035	0.2436	29.866
5770	1996 Nissan Sentra	Gasohol + AAE00, Seq#6	0.096	0.793	0.135	294.61	0.01	0.0848	28.814
5769	1994 Nissan Sentra	Gasohol + AAE00, Seq#6	0.17	1.846	0.381	314.49	0.019	0.1492	26.842
	Average		0.196	1.768	0.279	413.998	0.029	0.163	22.282
	Deviation +/-		-1.25%	-11.25%	-7.24%	-0.44%	-7.22%	0.16%	0.45%

INTERNATIONAL SEARCH REPORT

International Application No.

PCT/GB 99/04155

A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 C10L1/02 C10L1/14

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 C10L

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	WO 98 17745 A (WILLIAMSON IAN VERNON ;HAZEL CLIFFORD JAMES (GB)) 30 April 1998 (1998-04-30) cited in the application page 2, paragraph 2; claims 1.52; examples 2.6,10 page 4	1,5-15, 18-21,23
A	EP 0 466 511 A (ETHYL PETROLEUM ADDITIVES INC) 15 January 1992 (1992-01-15) page 5 --- -/--	21

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

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Date of the actual completion of the international search

19 April 2000

Date of mailing of the international search report

04/05/2000

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INTERNATIONAL SEARCH REPORT

International Application No.

PCT/GB 99/04155

C (Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	<p> DATABASE WPI Section Ch, Week 199641 Derwent Publications Ltd., London, GB: Class D25, AN 1996-408376 XP002136066 & JP 08 198830 A (KAWAKEN FINE CHEM CO LTD), 6 August 1996 (1996-08-06) abstract ----- </p>	16.17

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/GB 99/04155

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